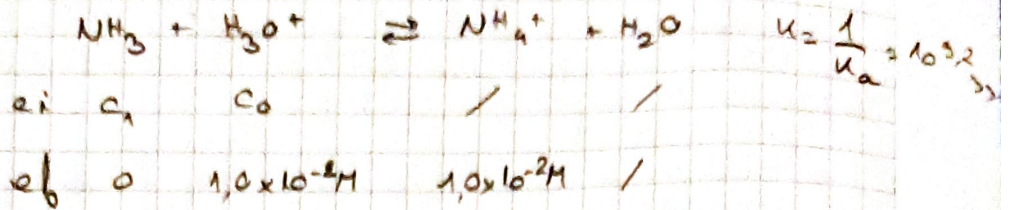
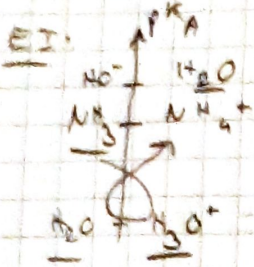
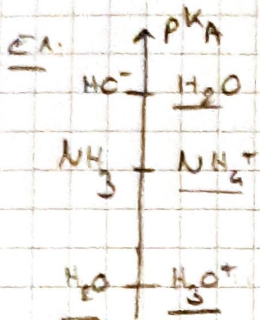


SAUCE 1 : ACIDE - BASE

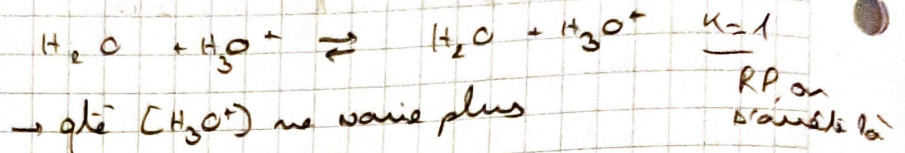
Application 7:



reaction totale

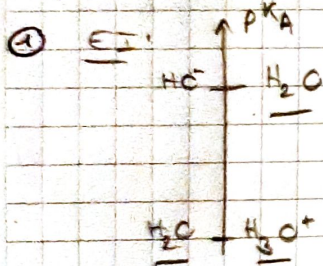


meilleure acide, meilleure base: H₂O, H₃O⁺

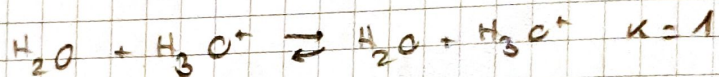


Donc $pH = -\log(10^{-2}) = 2$

Application 8:

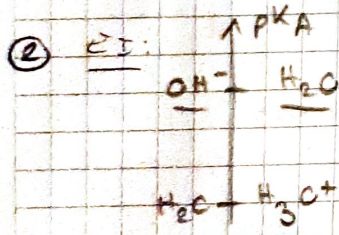


meilleure acide, meilleure base: H₂O, H₃O⁺

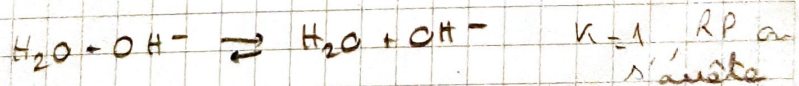


donc qte [H₃O⁺] ne varie plus

Donc $pH = -\log([H_3O^+]) = 3$

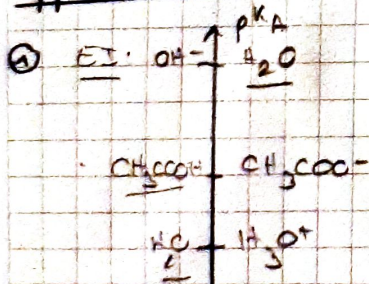


meilleure acide, meilleure base: H₂O, OH⁻

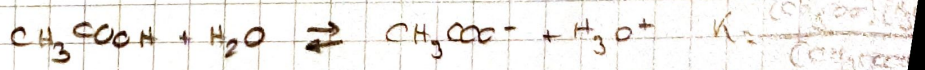


Donc $pH = 14 - pOH = 14 - 3 = 11$

Application 9



meilleure acide, meilleure base: CH₃COOH, H₂O



ei	10^{-2}	/	/	/
ef	10^{-2}	/	$10^{-3,4}$	$10^{-3,4}$

$1 \cdot 10^{-4,8} = \frac{x^2}{10^{-2}} \Rightarrow x^2 = 10^{-6,8} \Rightarrow x = 10^{-3,4} < 10^{-3}$ de

→ on suppose une RP peu avancée